



0,0651

-x

+x +x

0,0651-x

+x +x

$$K_b = \frac{K_w}{K_a} = \frac{1,00 \cdot 10^{-14}}{2,1 \cdot 10^{-4}} = 4,76 \cdot 10^{-11}$$

$$K_b = \frac{[\text{HCOOH}][\text{OH}^-]}{[\text{HCOO}^-]}$$

$$4,76 \cdot 10^{-11} = \frac{x^2}{0,0651 - x}$$

x transcurrible

$$x = 1,76 \cdot 10^{-6}$$

$$\text{pOH} = -\log 1,76 \cdot 10^{-6} = 5,75$$

$$\text{pH} = 14 - 5,75 = 8,25$$